Exam ID: 103

National Level ScienceX Olympiads

ScienceX Chemistry Olympiad (SCO)

Name: _____

Date: _____

Exam Instructions

1. The timer has been set for the exam, and a countdown will display the remaining time. When the timer runs out, the exam will end automatically. No further action will be required after that.
2. You can flag any question that you want.
3. If you don't want to submit any of your provided answers, you can clear them.
4. If you choose to drop this exam, it will not be submitted, and no result will be generated.
5. Once you finish the exam, you cannot resume it.

Secure Exam Advice

This exam is specially secured. Which means you cannot leave your browser window once the exam starts.

The following actions are not supported during your exam.

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- Taking a screenshot
- Pressing Ctrl + C
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- Pressing F12

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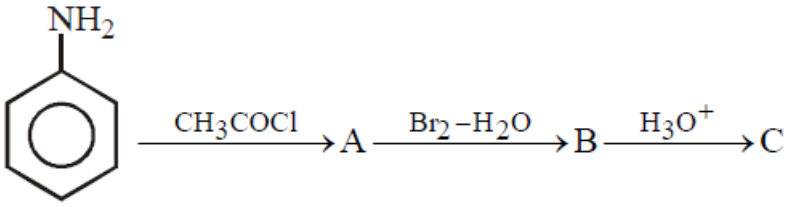
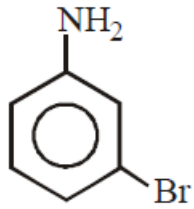
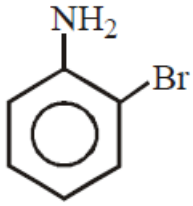


C (major product) is –

(a) 

(b) 

(c) 

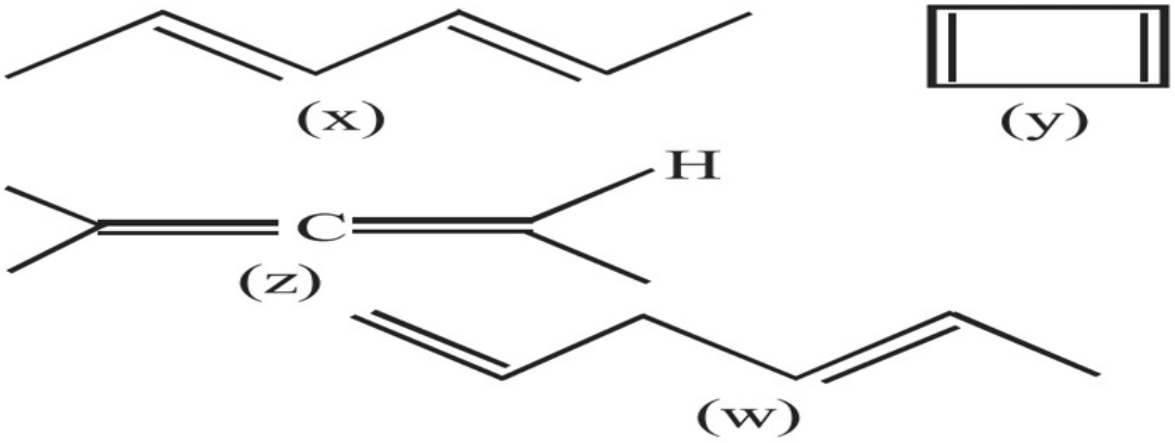
(d) None of these

Please mark (✓) for the correct answer.

- **A.** (b)
- **B.** (a)
- **C.** (d)
- **D.** (c)

Question: 2 of 50 QID: 222 Marks: 4

The correct stability order of following species is



Please mark (✓) for the correct answer.

- **A.** x > y > w > z
- **B.** z > x > y > w
- **C.** x > w > z > y
- **D.** y > x > w > z

Question: 3 of 50 QID: 231 Marks: 4

Nitrogen in an organic compound can be estimated by

Please mark (✓) for the correct answer.

- **A.** Neither (a) nor (b)
- **B.** Both (a) and (b)
- **C.** Duma's method only
- **D.** Kjeldahl's method only

Question: 4 of 50 QID: 213 Marks: 4

Assertion: The molecules of the dissolved gas present in a liquid gain kinetic energy as temperature is raised.

Reason: Gases tends to be more soluble in liquids as the temperature is raised

Please mark (✓) for the correct answer.

- **A.** If both Assertion and Reason are correct, but Reason is not the correct explanation of Assertion.
- **B.** If both Assertion and Reason are correct and Reason is the correct explanation of Assertion.
- **C.** If both the Assertion and Reason are incorrect.
- **D.** If Assertion is correct but Reason is incorrect.

Question: 5 of 50 QID: 218 Marks: 4

Assertion: Sn^{4+} compounds are stronger oxidizing agents than Pb^{4+} compounds.

Reason: The higher oxidation states for the group 14 elements are more stable for the heavier members of the group due to 'inert pair effect'.

Please mark (✓) for the correct answer.

- **A.** If both Assertion and Reason are correct and Reason is the correct explanation of Assertion.
- **B.** If Assertion is correct but Reason is incorrect.
- **C.** If both the Assertion and Reason are incorrect.
- **D.** If both Assertion and Reason are correct, but Reason is not the correct explanation of Assertion.

Question: 6 of 50 QID: 230 Marks: 4

Which of the following structure is similar to graphite?

Please mark (✓) for the correct answer.

- **A.** BN
- **B.** B_2H_6
- **C.** B_4C
- **D.** B

Question: 7 of 50 QID: 250 Marks: 4

Electron affinity of sulphur is

Please mark (✓) for the correct answer.

- **A.** less than O but more than Se
- **B.** equal to O and Se
- **C.** more than O but less than Se
- **D.** more than O and Se

Question: 8 of 50 QID: 229 Marks: 4

A compound X, of boron reacts with NH_3 on heating to give another compound Y which is called inorganic benzene. The compound X can be prepared by treating BF_3 with lithium aluminium hydride. The compounds X and Y are represented by the formulas.

Please mark (✓) for the correct answer.

- **A.** B_2H_6, $B_3N_3H_6$
- **B.** $B_3N_3H_6$, B_2H_6
- **C.** BF_3, $B_3N_3H_6$
- **D.** B_2O_3, $B_3N_3H_6$

Question: 9 of 50 QID: 239 Marks: 4

If salt bridge is removed from two half-cells the voltage

Please mark (✓) for the correct answer.

- **A.** increases gradually
- **B.** does not change
- **C.** drops to zero
- **D.** increases rapidly

Question: 10 of 50 QID: 207 Marks: 4

A gas cylinder with cooling gas inside can sustain pressures up to 14.9 atmospheres. At 27°C, the cylinder's pressure gauge reads 12 atmospheres. The temperature rises as a result of the building's unexpected fire. The cylinder explodes at the following temperature:

Please mark (✓) for the correct answer.

- **A.** 87.5°C
- **B.** 115.5°C
- **C.** 99.5°C
- **D.** 135.5°C

Question: 11 of 50 QID: 240 Marks: 4

What is the activation energy for a reaction if its rate doubles when the temperature is raised from 20°C to 35°C? ($R = 8.314\ J\ mol^{-1}\ K^{-1}$)

Please mark (✓) for the correct answer.

- **A.** 342 kJ mol^{-1}
- **B.** 15.1 kJ mol^{-1}
- **C.** 34.7 kJ mol^{-1}
- **D.** 269 kJ mol^{-1}

Question: 12 of 50 QID: 236 Marks: 4

Blood cells retain their normal shape in solution which are

Please mark (✓) for the correct answer.

- **A.** isotonic to blood
- **B.** equinormal to blood.
- **C.** hypertonic to blood
- **D.** hypotonic to blood

Question: 13 of 50 QID: 219 Marks: 4

Which of these does not happen when a moderate amount of sodium metal dissolves in liquid NH_3 at a low temperature?

Please mark (✓) for the correct answer.

- **A.** A blue-coloured solution is produced.
- **B.** Liquid ammonia maintains its diamagnetic properties
- **C.** Na^+ ions start to form in the mixture.
- **D.** Liquid NH_3 becomes an effective electrical conductor.

Question: 14 of 50 QID: 246 Marks: 4

Which of the following statements is correct ?

Please mark (✓) for the correct answer.

- **A.** Fluxes are carefully chosen to combine with the gangue present in the ore to produce easily fusible slag to carry away the impurities
- **B.** Gangues are carefully chosen to combine with the slag present in the ore to produce easily fusible flux to carry away the impurities
- **C.** Slags are carefully chosen to combine with the flux present in the ore to produce easily fusible gangue to carry away the impurities
- **D.** Gangues are carefully chosen to combine with the flux present in the ore to produce easily fusible slag to carry away the impurities

Question: 15 of 50 QID: 215 Marks: 4

The pH of solution containing 0.10 M sodium acetate and 0.03 M acetic acid is (pKa for CH_3COOH = 4.57)

Please mark (✓) for the correct answer.

- **A.** 6.09
- **B.** 4.09
- **C.** 7.09
- **D.** 5.09

Question: 16 of 50 QID: 225 Marks: 4

The first ionization energies of alkaline earth metals are higher than those of alkali metals. This is because

Please mark (✓) for the correct answer.

- **A.** there is increase in the nuclear charge of alkaline earth metals
- **B.** there is no change in the nuclear charge
- **C.** there is decrease in the nuclear charge of alkaline earth metals
- **D.** none of these

Question: 17 of 50 QID: 214 Marks: 4

Out of the following statements, which statement is incorrect?

Please mark (✓) for the correct answer.

- **A.** Heat of formation is defined as the heat exchange when one mole of a compound is formed from its constituent elements at STP.
- **B.** For all isothermal process change in internal energy is positive.
- **C.** Enthalpy change of a reaction does not depend upon different intermediate reactions.
- **D.** The internal energy of a substance increases with increase in temperature due to increase in rotational, translational and vibrational energy of the molecule.

Question: 18 of 50 QID: 210 Marks: 4

Out of the following statements, which statement (s) are correct?

(i) Ice melts faster at high altitude

(ii) Pressure cooker reduces cooking time because increase of pressure increases Boiling point.

(iii) Wet air is heavier than dry air.

(iv) All molecules in a gas have same speed.

Please mark (✓) for the correct answer.

- **A.** (ii)
- **B.** (ii) and (iv)
- **C.** (i) and (iii)
- **D.** (i)

Question: 19 of 50 QID: 209 Marks: 4

Based on equation $E = -2.178 \times 10^{-18} J(\frac{Z^2}{n^2})$, certain conclusions are written. Which of them is not correct?

Please mark (✓) for the correct answer.

- **A.** The negative sign in equation simply means that the energy or electron bound to the nucleus is lower than it would be if the electrons were at the infinite distance from the nucleus.
- **B.** Larger the value of n, the larger is the orbit radius.
- **C.** For n = 1, the electron has a more negative energy than it does for n = 6 which mean that the electron is more loosely bound in the smallest allowed orbit.
- **D.** Equation can be used to calculate the change in energy when the electron changes orbit.

Question: 20 of 50 QID: 245 Marks: 4

Which of the following statements related to Ellingham diagrams are correct ?

(i) It provides a sound basis for the choice of reducing agent in the reduction of oxides.

(ii) Each Ellingham plot is represented by a straight line untill unless there is some change in phase i.e. solid →liquid, liquid → gas and gas → liquid occurs.

(iii) Diagrams similar to Ellingham can be constructed for sulphides and halides which clearly indicates why reduction of M_xS is difficult in comparison to M_xO.

(iv) Ellingham diagrams predicts the tendency of reduction with a reducing agent and kinetics of the reduction process.

Please mark (✓) for the correct answer.

- **A.** (ii) and (iv)
- **B.** (i), (ii) and (iii)
- **C.** (i), (ii) and (iv)
- **D.** (i) and (iii)

Question: 21 of 50 QID: 249 Marks: 4

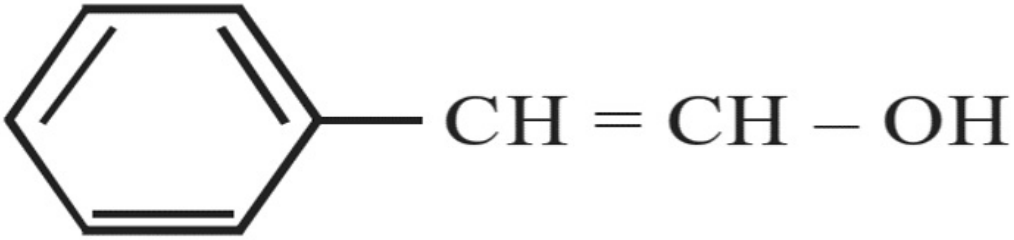
Which of the following is incorrect?

Please mark (✓) for the correct answer.

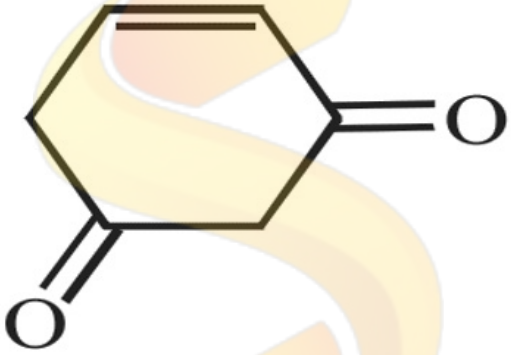
- **A.** Cu does not shows +2 oxidation state with I^-.
- **B.** V shows oxidation state of + 5 in VF_5.
- **C.** Fe and Co shows +3 oxidation state in FeX_3 and CoF_3.
- **D.** Mn shows oxidation state of +7 in MnF_7

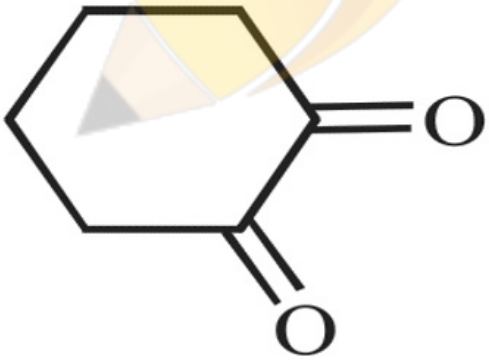
Question: 22 of 50 QID: 223 Marks: 4

Tautomerism is exhibited by

(1) 

(2) 

(3) 

(4) 

Please mark (✓) for the correct answer.

- **A.** (2), (3), and (4)
- **B.** All of these
- **C.** None of these
- **D.** (1), (3) and (4)

Question: 23 of 50 QID: 243 Marks: 4

Read the following statements
(i) Tyndall effect is used to distinguish between a colloidal and true solution.
(ii) Values of colligative properties are same for true and colloidal solutions.
(iii) Random bombardment of the colloidal particles by the molecules of the dispersion medium does not allow colloids to settle thereby providing stability to them.
(iv) Most acceptable phenomena to account for the charge of sol particles is electrodispersion.

Which of the following is the correct code for statements above?

Please mark (✓) for the correct answer.

- **A.** FTFT
- **B.** TTFF
- **C.** TFTF
- **D.** TFFT

Question: 24 of 50 QID: 216 Marks: 4

In which of the following acid-base titration, pH is greater than 8 at equivalence point.

Please mark (✓) for the correct answer.

- **A.** Hydrochloric acid versus ammonia
- **B.** Acetic acid versus ammonia
- **C.** Hydrochloric acid versus sodium hydroxide
- **D.** Acetic acid versus sodium hydroxide

Question: 25 of 50 QID: 252 Marks: 4

In the commercial manufacture of ethyl alcohol from starchy substances by fermentation method. Which enzymes slipwise complete the fermentation reaction

Please mark (✓) for the correct answer.

- **A.** Diastase, invertase and zymase
- **B.** Diastase, zymase and lactase
- **C.** Maltase, zymase and invertase
- **D.** Diastase, maltase and zymase

Question: 26 of 50 QID: 205 Marks: 4

Match List I and List II and identify the correct matching codes from the given choices

List I Compound	List II Structure
A. ClF_3	1. Square planar
B: PCl_5	2. Tetrahedral
C: IF_5	3. Trigonal bipyramidal
D: CCl_4	4. Square pyramidal
E: XeF_4	5. T-shaped

Please mark (✓) for the correct answer.

- **A.** A-5, B-4, C-3, D-2, E-1
- **B.** A-4, B-3, C-5, D-2, E-1
- **C.** A-5, B-3, C-4, D-1, E-2
- **D.** A-5, B-3, C-4, D-2, E-1

Question: 27 of 50 QID: 241 Marks: 4

The bromination of acetone that occurs in acid solution is represented by this equation.
$CH_3COCH_3(aq) + Br_2(aq) \rightarrow CH_3COCH_2Br(aq) + H^+(aq) + Br^-(aq)$
These kinetic data were obtained for given reaction concentrations.

Initial Concentrations, M			Initial rate, disappearance of Br_2, Ms^{-1}
$[CH_3COCH_3]$	$[Br_2]$	$[H^+]$	
0.30	0.05	0.05	5.7×10^{-5}
0.30	0.10	0.05	5.7×10^{-5}
0.30	0.10	0.10	1.2×10^{-4}
0.40	0.05	0.20	3.1×10^{-4}

Based on given data, the rate equations is:

Please mark (✓) for the correct answer.

- **A.** Rate = k $[CH_3COCH_3][Br_2]$
- **B.** Rate = k $[CH_3COCH_3][Br_2]$ $[H^+]$
- **C.** Rate = k $[CH_3COCH_3]$ $[Br_2]$ $[H^+]^2$
- **D.** Rate = k$[CH_3COCH_3][H^+]$

Question: 28 of 50 QID: 220 Marks: 4

Select correct statement(s).

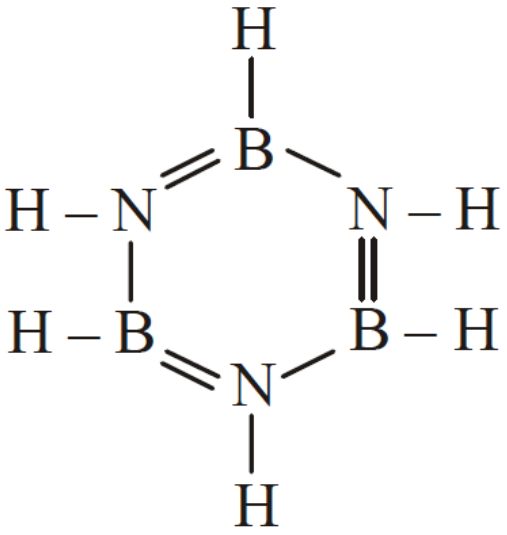
(i) Cyanamide ion (CN_2^{2-}) is isoelectronic with CO_2 and has the same linear structure.
(ii) Mg_2C_3 reacts with water to form propyne
(iii) CaC_2 has NaCl type lattice.
(iv) Chemically borax is sodium tetraborate decahydrate

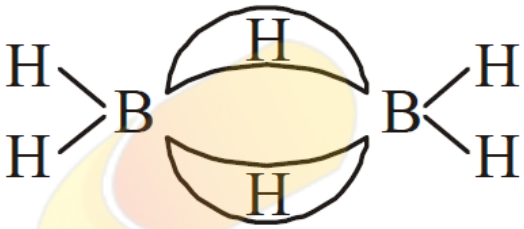
Please mark (✓) for the correct answer.

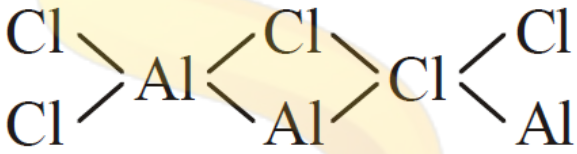
- **A.** (i), (ii) and (iii)
- **B.** (i), (ii), (iii) and (iv)
- **C.** (i) and(ii)
- **D.** (i) and (iii)

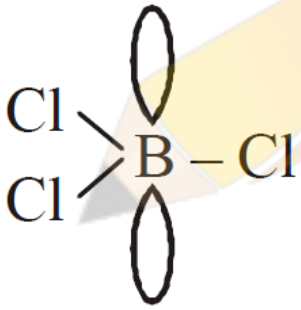
Question: 29 of 50 QID: 228 Marks: 4

Which of the following compounds is not matched correctly with its structure?

(a)  – Borazine

(b)  – Diborane

(c)  – Aluminium chloride

(d)  – Boron trichloride

Please mark (✓) for the correct answer.

- **A.** (b)
- **B.** (c)
- **C.** (d)
- **D.** (a)

Question: 30 of 50 QID: 247 Marks: 4

Larger number of oxidation states are exhibited by the actinoids than those by the lanthanoids, the main reason being

Please mark (✓) for the correct answer.

- **A.** 4f orbitals more diffused than the 5f orbitals
- **B.** more energy difference between 5f and 6d than between 4f and 5d orbitals
- **C.** more reactive nature of the actionids than the lanthanoids
- **D.** lesser energy difference between 5f and 6d than between 4f and 5d orbitals

Question: 31 of 50 QID: 221 Marks: 4

The end product (C) in the following sequence of reactions is

$$HC \equiv CH \xrightarrow[20\%\ H_2SO_4]{1\%\ HgSO_4} A \xrightarrow[H_2O]{CH_3MgX} B \xrightarrow{[O]} (C)$$

Please mark (✓) for the correct answer.

- **A.** ethanol
- **B.** acetone
- **C.** acetic acid
- **D.** isopropyl alcohol

Question: 32 of 50 QID: 233 Marks: 4

Which of the following statement(s) is/are correct ?
(i) Classical smog is a mixture of smoke, fog and sulphur dioxide.
(ii) Classical smog is also called oxidising smog
(iii) Hydrocarbons, NO2 and PAN are components of photochemical smog.

Please mark (✓) for the correct answer.

- **A.** (i), (ii) and (iii)
- **B.** (i) and (iii)
- **C.** (iii) only
- **D.** (i) and (ii)

Question: 33 of 50 QID: 238 Marks: 4

E° of a cell $aA + bB \longrightarrow cC + dD$ is

(a) $E + RT \ln \dfrac{[a]^A[b]^B}{[c]^C[d]^D}$ (b) $E + \dfrac{RT}{nF} \ln \dfrac{[C]^c[D]^d}{[A]^a[B]^b}$

(c) $E + \dfrac{RT}{nF} \ln \dfrac{[C]^c[d]^D}{[A]^A[B]^B}$ (d) $E + \dfrac{RT}{nF} \ln \dfrac{[a]^A[B]^B}{[C]^C[d]^D}$

Please mark (✓) for the correct answer.

- **A.** (c)
- **B.** (b)
- **C.** (a)
- **D.** (d)

Question: 34 of 50 QID: 211 Marks: 4

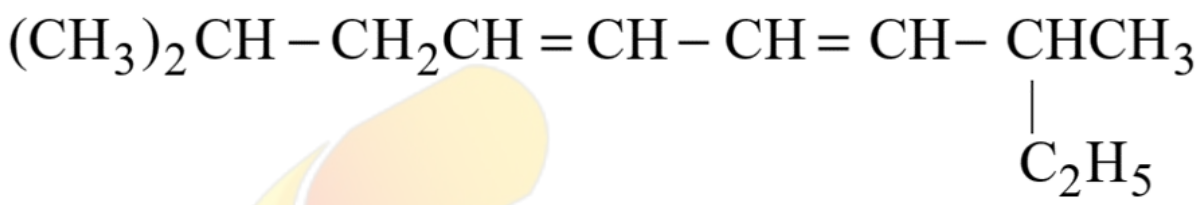
1 g equimolecular mixture of Na_2CO_3 and $NaHCO_3$ is reacted with 0.1 N HCl. The mL of 0.1N HCl required to react completely with the above mixture is:

Please mark (✓) for the correct answer.

- **A.** 153.8 mL
- **B.** 198.4 mL
- **C.** 15.78 mL
- **D.** 295.5 mL

Question: 35 of 50 QID: 224 Marks: 4

The IUPAC name of the following compound is



Please mark (✓) for the correct answer.

- **A.** 2,8-dimethyl-4,6-decadiene
- **B.** 2,8-dimethyl-3,6-decadiene
- **C.** 1,1,7,7-tetramethyl-2,5-octadiene
- **D.** 1,5-di-iso-propyl-1,4-hexadiene

Question: 36 of 50 QID: 226 Marks: 4

Which of the following sequence of T and F is correct for given statements?
(i) The alkali metal hydroxides are the strongest of all bases.
(ii) All alkali metal halides have high negative enthalpies of formation.
(iii) The stability of the carbonates and hydrogen carbonates of alkali metals decrease with increase in electropositive character down the group.
(iv) Only $LiHCO_3$ exist as solid.

Please mark (✓) for the correct answer.

- **A.** TTFF
- **B.** TFFT
- **C.** FTFT
- **D.** TTTT

Question: 37 of 50 QID: 208 Marks: 4

The amount of O_2 released during the electrolysis of water is 2.24 dm^3. Under the same circumstances, the amount of hydrogen released will be

Please mark (✓) for the correct answer.

- **A.** 2.24 dm^3
- **B.** 4.48 dm^3
- **C.** 0.56 dm^3
- **D.** 1.12 dm^3

Question: 38 of 50 QID: 251 Marks: 4

Which of the following statements are correct?

(i) As a result of lanthanoid contraction members of 4d and 5d series exhibit similar radii.

(ii) IE_2 is high for Cr and Cu whereas IE_3 is very high for Zn.

(iii) Heavier members of d-block elements like p-block elements favours lower oxidation states.

(iv) In any transition series maximum number of oxidation states is shown by middle elements or elements near middle elements.

Please mark (✓) for the correct answer.

- **A.** (i), (ii) and (iv)
- **B.** (i) and (ii)
- **C.** (i), (ii) and (iii)
- **D.** (ii) and (iv)

Question: 39 of 50 QID: 242 Marks: 4

The adsorption of a gas on a solid surface varies with pressure of the gas in which of the following manner

Please mark (✓) for the correct answer.

- **A.** Slow →fast → independent of the pressure
- **B.** Independent of the pressure → slow → fast
- **C.** Fast → slow → independent of the pressure
- **D.** Independent of the pressure → fast → slow

Question: 40 of 50 QID: 248 Marks: 4

In nitrogen family, the H-M-H bond angle in the hydridesgradually becomes closer to 90º on going from N to Sb. This shows that gradually

Please mark (✓) for the correct answer.

- **A.** The basic strength of the hydrides increases
- **B.** Almost pure p-orbitals are used for M-H bonding
- **C.** The bond energies of M-H bonds increase
- **D.** The bond pairs of electrons become nearer to the central atom

Question: 41 of 50 QID: 212 Marks: 4

Match List – I with List –II for the compositions of substances and select the correct answer using the code given below the lists

List - I Substances	List - II Composition
(A) Plaster of paris	(i) $CaSO_4.2H_2O$
(B) Epsomite	(ii) $CaSO_4.\frac{1}{2}H_2O$
(C) Kieserite	(iii) $MaSO_4.7H_2O$
(D) Gypsum	(iv) $MgSO_4.H_2O$

Please mark (✓) for the correct answer.

- **A.** (A) →(ii), (B) →(iii), (C) →(iv), (D) →(i)
- **B.** (A) →(iii), (B) →(iv), (C) →(i), (D) →(ii)
- **C.** (A) →(ii), (B) →(iii), (C) →(i), (D) →(iv)
- **D.** (A) →(iii), (B) →(iv), (C) →(ii), (D) →(i)

Question: 42 of 50 QID: 253 Marks: 4

The IUPAC name of the complex $[Co(NH_3)_4(H_2O)Cl]Cl_2$ is

Please mark (✓) for the correct answer.

- **A.** aquatetramminechloridocobalt (III) chloride
- **B.** chloridoaquatetramminechloridocobalt (III) chloride
- **C.** tetrammineaquachloridocobalt (III) chloride
- **D.** chloridoaquatetramminechloridocobalt (III) chloride

Question: 43 of 50 QID: 235 Marks: 4

Which of the following sequence of T and F is true for given statements. Here T stands for true statement and F stands for false statement?
(i) Frenkel defect results in increase in density of the solid.
(ii) ZnS, AgCl, AgBr and AgI shows Frenkel defect.
(iii) Schottky defect results in decrease in density of the solid.
(iv) AgBr shows Schottky defect only.
(v) For NaCl there is one Schottky defect per 10^{16} ions.

Please mark (✓) for the correct answer.

- **A.** FTTFT
- **B.** TTTFT
- **C.** TTTFF
- **D.** FTTFF

Question: 44 of 50 QID: 217 Marks: 4

Identify the incorrect statement

Please mark (✓) for the correct answer.

- **A.** Water gas is produced by passing steam over hot coke.
- **B.** Dry ice evaporates at −78°C without melting.
- **C.** Producer gas is a mixture of CO and N_2.
- **D.** Bell metal is an alloy of Cu and Zn

Question: 45 of 50 QID: 227 Marks: 4

Which of the following on thermal decomposition yields a basic as well as acidic oxide ?

Please mark (✓) for the correct answer.

- **A.** $CaCO_3$
- **B.** $NaNO_3$
- **C.** NH_4NO_3
- **D.** $KClO_3$

Question: 46 of 50 QID: 244 Marks: 4

Which of the following does not contain a hydrophobic structure ?

Please mark (✓) for the correct answer.

- **A.** Lanolin
- **B.** Rubber
- **C.** Glycogen
- **D.** Linseed oil

Question: 47 of 50 QID: 234 Marks: 4

Assertion : Suspended particulate matter (SPM) is an important pollutant released by diesel vehicles.

Reason : Catalytic converters greatly reduce pollution caused by automobiles.

Please mark (✓) for the correct answer.

- **A.** Assertion is correct, reason is correct; reason is not a correct explanation for assertion
- **B.** Assertion is incorrect, reason is correct
- **C.** Assertion is correct, reason is correct; reason is a correct explanation for assertion.
- **D.** Assertion is correct, reason is incorrect

Question: 48 of 50 QID: 206 Marks: 4

Out of the following statements regarding ionization potential, which one is true?

Please mark (✓) for the correct answer.

- **A.** It decreases with an increase in atomic radii
- **B.** It is independent of atomic radii
- **C.** It increases with an increase in atomic radii
- **D.** It remains constant with change in atomic radii

Question: 49 of 50 QID: 232 Marks: 4

Which of the following sequence of T and F is correct for given statements. Here T stands for correct and F stands for false statement
(i) The more the number of contributing structures, the more is the resonance energy.
(ii) The resonance structures have different positions of nuclei but same number of unpaired electrons
(iii) The energy of actual structure of the molecule (the resonance hybrid) is lower than that of any of the canonical structures.

Please mark (✓) for the correct answer.

- **A.** FFF
- **B.** TFF
- **C.** TTT
- **D.** TFT

Question: 50 of 50 QID: 237 Marks: 4

At 300 K the vapour pressure of an ideal solution containing 1 mole of liquid A and 2 moles of liquid B is 500 mm of Hg. The vapour pressure of the solution increases by 25 mm of Hg, if one more mole of B is added to the above ideal solution at 300 K. Then the vapour pressure of A in its pure state is

Please mark (✓) for the correct answer.

- **A.** 400 mm of Hg
- **B.** 300 mm of Hg
- **C.** 600 mm of Hg
- **D.** 500 mm of Hg

--- END OF QUESTION PAPER ---

Answer Key

No	Question Type	QID	Correct Answer
Question - 1	Multiple Choice (Radiobutton)	254	D
Question - 2	Multiple Choice (Radiobutton)	222	C
Question - 3	Multiple Choice (Radiobutton)	231	B
Question - 4	Multiple Choice (Radiobutton)	213	D
Question - 5	Multiple Choice (Radiobutton)	218	C
Question - 6	Multiple Choice (Radiobutton)	230	A
Question - 7	Multiple Choice (Radiobutton)	250	D
Question - 8	Multiple Choice (Radiobutton)	229	A
Question - 9	Multiple Choice (Radiobutton)	239	C
Question - 10	Multiple Choice (Radiobutton)	207	C
Question - 11	Multiple Choice (Radiobutton)	240	C
Question - 12	Multiple Choice (Radiobutton)	236	A
Question - 13	Multiple Choice (Radiobutton)	219	B
Question - 14	Multiple Choice (Radiobutton)	246	A
Question - 15	Multiple Choice (Radiobutton)	215	D
Question - 16	Multiple Choice (Radiobutton)	225	A
Question - 17	Multiple Choice (Radiobutton)	214	B
Question - 18	Multiple Choice (Radiobutton)	210	A
Question - 19	Multiple Choice (Radiobutton)	209	C
Question - 20	Multiple Choice (Radiobutton)	245	D
Question - 21	Multiple Choice (Radiobutton)	249	D
Question - 22	Multiple Choice (Radiobutton)	223	D
Question - 23	Multiple Choice (Radiobutton)	243	C
Question - 24	Multiple Choice (Radiobutton)	216	D
Question - 25	Multiple Choice (Radiobutton)	252	D
Question - 26	Multiple Choice (Radiobutton)	205	D
Question - 27	Multiple Choice (Radiobutton)	241	D
Question - 28	Multiple Choice (Radiobutton)	220	B
Question - 29	Multiple Choice (Radiobutton)	228	B
Question - 30	Multiple Choice (Radiobutton)	247	D
Question - 31	Multiple Choice (Radiobutton)	221	B
Question - 32	Multiple Choice (Radiobutton)	233	B
Question - 33	Multiple Choice (Radiobutton)	238	B
Question - 34	Multiple Choice (Radiobutton)	211	A
Question - 35	Multiple Choice (Radiobutton)	224	A
Question - 36	Multiple Choice (Radiobutton)	226	A
Question - 37	Multiple Choice (Radiobutton)	208	B
Question - 38	Multiple Choice (Radiobutton)	251	A
Question - 39	Multiple Choice (Radiobutton)	242	C
Question - 40	Multiple Choice (Radiobutton)	248	B
Question - 41	Multiple Choice (Radiobutton)	212	A
Question - 42	Multiple Choice (Radiobutton)	253	C
Question - 43	Multiple Choice (Radiobutton)	235	A

No	Question Type	QID	Correct Answer
Question - 44	Multiple Choice (Radiobutton)	217	D
Question - 45	Multiple Choice (Radiobutton)	227	A
Question - 46	Multiple Choice (Radiobutton)	244	B
Question - 47	Multiple Choice (Radiobutton)	234	A
Question - 48	Multiple Choice (Radiobutton)	206	A
Question - 49	Multiple Choice (Radiobutton)	232	D
Question - 50	Multiple Choice (Radiobutton)	237	B

--- END OF ANSWER KEY ---